



This is “Appendix D: Dissociation Constants and pK_b Values for Bases at 25°C”, appendix 4 from the book [Principles of General Chemistry \(index.html\)](#) (v. 1.0M).

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Chapter 28

Appendix D: Dissociation Constants and pK_b Values for Bases at 25°C

Name	Formula	K_b	pK_b
Ammonia	NH_3	1.8×10^{-5}	4.75
Aniline	$C_6H_5NH_2$	7.4×10^{-10}	9.13
<i>n</i> -Butylamine	$C_4H_9NH_2$	4.0×10^{-4}	3.40
<i>sec</i> -Butylamine	$(CH_3)_2CHCH_2NH_2$	3.6×10^{-4}	3.44
<i>tert</i> -Butylamine	$(CH_3)_3CNH_2$	4.8×10^{-4}	3.32
Dimethylamine	$(CH_3)_2NH$	5.4×10^{-4}	3.27
Ethylamine	$C_2H_5NH_2$	4.5×10^{-4}	3.35
Hydrazine	N_2H_4	1.3×10^{-6}	5.9
Hydroxylamine	NH_2OH	8.7×10^{-9}	8.06
Methylamine	CH_3NH_2	4.6×10^{-4}	3.34
Propylamine	$C_3H_7NH_2$	3.5×10^{-4}	3.46
Pyridine	C_5H_5N	1.7×10^{-9}	8.77
Trimethylamine	$(CH_3)_3N$	6.3×10^{-5}	4.20

Source of data: *CRC Handbook of Chemistry and Physics*, 84th Edition (2004).